



THE UNIVERSITY *of* EDINBURGH

## Edinburgh Research Explorer

### Dihydrogen bonds in solid BH<sub>3</sub>NH<sub>3</sub>

**Citation for published version:**

Morrison, C & Siddick, MM 2004, 'Dihydrogen bonds in solid BH<sub>3</sub>NH<sub>3</sub>', *Angewandte Chemie International Edition*, vol. 108, no. 1, pp. 4780-4782. <https://doi.org/10.1002/anie.200460096>

**Digital Object Identifier (DOI):**

[10.1002/anie.200460096](https://doi.org/10.1002/anie.200460096)

**Link:**

[Link to publication record in Edinburgh Research Explorer](#)

**Document Version:**

Peer reviewed version

**Published In:**

Angewandte Chemie International Edition

**Publisher Rights Statement:**

Copyright © 2004 WILEY-VCH Verlag GmbH & Co. KGaA, Weinheim. All rights reserved.

**General rights**

Copyright for the publications made accessible via the Edinburgh Research Explorer is retained by the author(s) and / or other copyright owners and it is a condition of accessing these publications that users recognise and abide by the legal requirements associated with these rights.

**Take down policy**

The University of Edinburgh has made every reasonable effort to ensure that Edinburgh Research Explorer content complies with UK legislation. If you believe that the public display of this file breaches copyright please contact [openaccess@ed.ac.uk](mailto:openaccess@ed.ac.uk) providing details, and we will remove access to the work immediately and investigate your claim.



This is the peer-reviewed version of the following article:

Morrison, C., & Siddick, M. M. (2004). Dihydrogen bonds in solid  $\text{BH}_3\text{NH}_3$ . *Angewandte Chemie International Edition*, 108 (1), 4780-4782.

which has been published in final form at <http://dx.doi.org/10.1002/anie.200460096>

This article may be used for non-commercial purposes in accordance with Wiley Terms and Conditions for self-archiving (<http://olabout.wiley.com/WileyCDA/Section/id-817011.html>).

Manuscript received: 23/03/2004; Revised: 28/05/2004; Article published: 07/09/2004

## Dihydrogen Bonds in Solid $\text{BH}_3\text{NH}_3$ \*\*

Carole A. Morrison<sup>1</sup> and Muhammad M. Siddick<sup>2</sup>

<sup>[1]</sup>EaStCHEM, School of Chemistry, Joseph Black Building, University of Edinburgh, West Mains Road, Edinburgh, EH9 3JJ, UK.

<sup>[2]</sup>School of Chemistry and School of Physics, University of Edinburgh, The King's Buildings, Mayfield Road, Edinburgh, EH9 3JZ, UK.

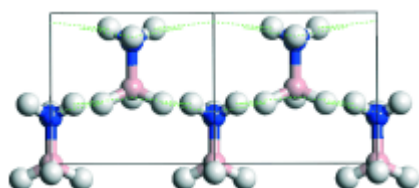
<sup>[\*]</sup>Corresponding author; e-mail: [c.morrison@ed.ac.uk](mailto:c.morrison@ed.ac.uk), fax: (+44) 131-650-4743

<sup>[\*\*]</sup>C.A.M. gratefully acknowledges The Royal Society for the award of a University Research Fellowship and the University of Edinburgh (School of Chemistry) for financial support. M.M.S. is grateful to the University of Edinburgh (Schools of Chemistry and Physics) for the award of a Research Studentship.

### Supporting information:

Supporting information for this article is available on the WWW under [http://www.wiley-vch.de/contents/jc\\_2002/2004/z460096\\_s.pdf](http://www.wiley-vch.de/contents/jc_2002/2004/z460096_s.pdf) or from the author.

### Graphical abstract:



### Keywords:

boron; density functional calculations; dihydrogen bonds; solid state structures

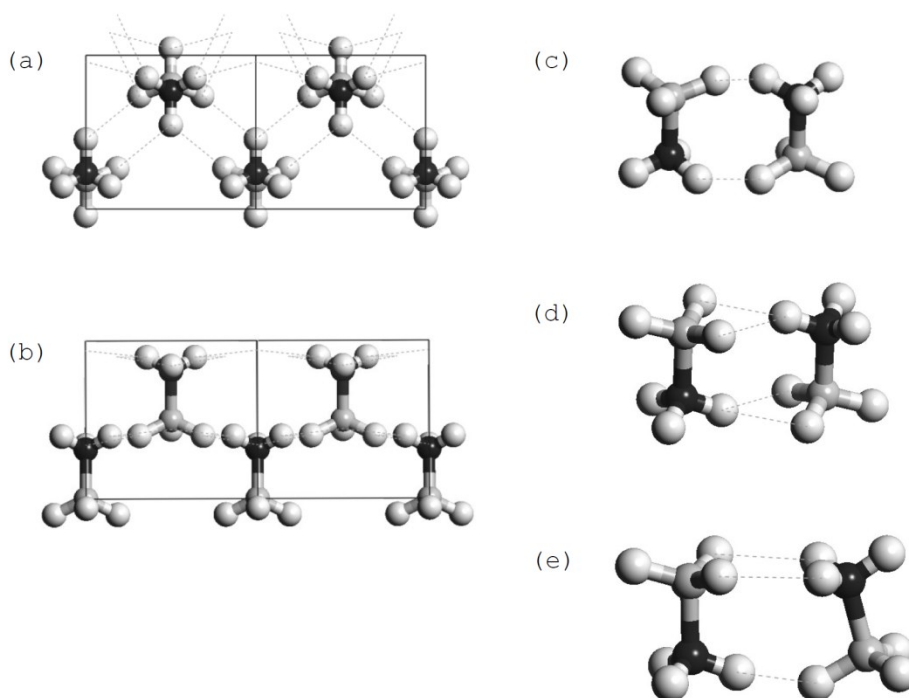
## Abstract

How strong is the H $\cdots$ H bond in solid BH<sub>3</sub>NH<sub>3</sub>? Insight from periodic density functional calculations predicts an average value of 12.7 kJ mol<sup>-1</sup>, reassigning the interaction towards the lower end of the dihydrogen bond strength spectrum.

## Main text

The “dihydrogen bond” D—H $\delta^+$  $\cdots\delta^-$ H—E—where D is a typical electrostatic atom such as nitrogen or oxygen, and E is a transition metal or boron—was first described by Crabtree and co-workers.<sup>1</sup> Both inter- and intramolecular versions of this bond have been cited, and with typical energies in the range of 12–28 kJ mol<sup>-1</sup>,<sup>2</sup> they are comparable in strength to conventional hydrogen bonds. In particular, the N—H $\delta^+$  $\cdots\delta^-$ H—B intermolecular contact in solid-state BH<sub>3</sub>NH<sub>3</sub> has received special attention.<sup>1–4</sup> The absence of nonbonding valence electrons on boron negates a possible H $\delta^+$  $\cdots\delta^-$ E interaction that could arise if E were a d-block element. Thus, for BH<sub>3</sub>NH<sub>3</sub> the H $\cdots$ H contact must result from a sigma-type interaction.

To date, all computational studies investigating properties of this unusual intermolecular bond refer to isolated dimer (i.e., gas-phase) models, which do not resemble the crystal structure all that closely (Figure 1 a, b1). The global minimum calculated by Crabtree and co-workers contains two H $\cdots$ H contacts (PCI-80/B3 LYP, Figure 1 c), which are each assigned an average bond strength of 25.5 kJ mol<sup>-1</sup> (no basis set super-position error (BSSE) or zero-point energy (ZPE) corrections applied).<sup>2</sup> This value has been described as being “surprisingly large” by the authors, and places the N—H $\cdots$ H—B contact at the upper end of the energy range quoted for dihydrogen bonds. Li et al. cite a different structure comprising four dihydrogen bonds (Figure 1 d), with an average bond strength of 11.6 kJ mol<sup>-1</sup> (after BSSE and ZPE correction; 6-31++G\*/MP2), that is, at the lower end of the dihydrogen bond energy spectrum.<sup>3</sup> Popelier characterized the intermolecular interaction on the basis of the calculated electron density for a third dimer containing three dihydrogen bonds (Figure 1 e), with an average bond energy of 18.3 kJ mol<sup>-1</sup> (although bond ellipticities indicate that not all bonds are of equal strength; 6-31G\*/MP2, no BSSE or ZPE correction applied).<sup>4</sup>



**Figure 1.** Crystal structure of  $\text{BH}_3\text{NH}_3$  showing the dihydrogen intermolecular bonding network as viewed along the  $c$  (a) and  $b$  axes (b). Dimer models used by Crabtree and co-workers. (c), Li et al. (d), and Popelier (e) to investigate properties of the  $\text{H}\cdots\text{H}$  contacts.

Given that the dimer models do not mimic the crystal structure all that closely, it is now timely to revisit the dihydrogen bond found in solid  $\text{BH}_3\text{NH}_3$  using a periodic quantum mechanical modeling approach, where the full crystallographic unit cell is used as the model for calculation. An optimized structure (unit cell and atomic positions) has been obtained with plane-wave density functional theory (PW-DFT) and the energy per molecule in the crystal lattice calculated. Deleting all but one molecule from the crystal lattice, and stretching the unit cell parameters (i.e., creating a pseudo-isolated molecule) effectively takes us to the gas phase. As the same basis set and level of theory are used in both calculations, comparison of the energies obtained gives the interaction energy, from which the average hydrogen bond strength can be obtained. (A full discussion of the computational methods is available in the Supporting Information.) We have recently demonstrated the success of this method for the test systems ammonia and urea.<sup>5</sup>

The results obtained from the PW-DFT calculations are given in Table 1. Although the simulation slightly overestimates the volume of the unit cell (+2 %), this is commonly observed with generalized gradient approximation (GGA) DFT functionals and translates to an increase in lattice parameters of no more than 0.1 Å. Similarly, all bond distances are reproduced to within approximately 0.1 Å. In the supercell calculation (i.e. pseudo-isolated molecule) the B—N bond lengthens by 0.044 Å, which is consistent with experimental values ( $\Delta 0.078$  Å in going from the solid state to the gas phase).

**Table 1.** Selected parameters for BH<sub>3</sub>NH<sub>3</sub>.

Parameters	Experimental		Calculated				
	solid <sup>[a]</sup>	gas <sup>[b]</sup>	solid	supercell			
				not relaxed	relaxed	dimer	monomer, relaxed
lattice							
<i>a</i> [Å]	5.395(2)	—	5.366	10.0	10.0	12.0	12.0
<i>b</i> [Å]	4.887(2)	—	4.914	10.0	10.0	12.0	12.0
<i>c</i> [Å]	4.986(2)	—	5.100	10.0	10.0	12.0	12.0
$\alpha, \beta, \gamma$ [°]	90	—	90	90	90	90.0	90.0
<i>Z</i>	2	—	2	1	1	2	1
volume [Å <sup>3</sup> ]	131.5(16)	—	134.5	1000.0	1000.0	1728	1728
space/point group	<i>Pmn</i> 2 <sub>1</sub>	—	<i>Pmn</i> 2 <sub>1</sub>	<i>P</i> 1	<i>P</i> 1	<i>C</i> <sub>2</sub> <i>h</i>	<i>C</i> <sub>3</sub> <i>v</i>
geometry							
<i>r</i> B—N [Å]	1.58(2)	1.6576(16)	1.580	1.580	1.624	1.606	1.624
<i>r</i> H...H(1) [Å]	2.02(3)	—	1.909	1.909	—	1.940	—
<i>r</i> H...H(2) [Å]	2.23(4)	—	2.174	2.174	—	—	—
<i>r</i> H...H(3) [Å]	2.21(4)	—	2.271	2.271	—	—	—
total energy [eV]	—	—	−896.43917	−447.39887	−447.43142	−895.50251	−447.41750
lattice energy [kJ mol <sup>−1</sup> ] <sup>[c]</sup>	—	—	—	79.2	—	—	—
sublimation energy [kJ mol <sup>−1</sup> ] <sup>[d]</sup>	—	—	—	—	76.0	—	—
H-bond strength [kJ mol <sup>−1</sup> ]	—	—	—	—	12.7 <sup>[e]</sup>	—	16.1 <sup>[f]</sup>

[a] Neutron diffraction study.<sup>7</sup> [b] Microwave spectroscopy study.<sup>8</sup> [c]  $\Delta H_{\text{lattice}}(0$

K)= $E_{\text{unit cell}}/Z - E_{\text{supercell (not relaxed)}}$ . [d]  $\Delta H_{\text{sub}}(0 \text{ K})=E_{\text{unit cell}}/Z - E_{\text{supercell (relaxed)}}$ . [e] H-bond

strength= $\Delta H_{\text{sub}}/(\text{intermolecular bond order}=6)$ . [f] H-bond strength= $[(E_{\text{dimer}}/2 - E_{\text{monomer (relaxed)}})/(\text{intermolecular bond order}=2)]$

Comparing the energy per molecule from the crystal lattice with that from the supercell calculation (with one molecule still in the crystal geometry, that is, no atom relaxation at this stage) enables the lattice energy<sup>6</sup> at 0 K for the crystal to be determined (79.2 kJ mol<sup>-1</sup>). Taking atomic relaxation for the isolated molecule into account gives a value for the sublimation energy<sup>[6]</sup> at 0 K of 76.0 kJ mol<sup>-1</sup>.<sup>§,¥</sup> As each BH<sub>3</sub>NH<sub>3</sub> molecule in the crystal lattice is connected by twelve intermolecular dihydrogen bonds to its neighbours [*i.e.* an intermolecular bond order of six, see Figure 1(a)] we thus estimate the average value for the dihydrogen bond present in the crystal structure to be of the order 12.7 kJ mol<sup>-1</sup>.<sup>⊕,∅</sup>

To demonstrate that the PW-DFT supercell method can reliably calculate binding energies for this system, we have also performed calculations for the dimer structure obtained by Li et al. (see Figure 1 d and Table 1) for direct comparison with the results obtained from ab initio molecular orbital calculations.<sup>3</sup> Comparing the energies of the dimer and monomer structures located in the center of a 12×12×12 Å supercell, we calculate the total binding energy (per molecule) to be 32.2 kJ mol<sup>-1</sup>, thus giving a dihydrogen bond strength for this complex of 16.1 kJ mol<sup>-1</sup>. This is largely consistent with Li's value of 13.4 kJ mol<sup>-1</sup> (after BSSE correction, ZPE correction not applied for consistency with our calculations). We note that our optimized structure exhibits a slightly shorter H...H interaction (1.94 versus 2.04 Å obtained by Li et al.), and this may explain why our simulations return a slightly stronger bond.

In conclusion, we have performed PW-DFT calculations to investigate the strength of the dihydrogen bond in the solid-state structure of BH<sub>3</sub>NH<sub>3</sub>. The initial parameters used in our periodic calculations are taken directly from the experimental crystallographic coordinates and unit cell, thus bypassing the need to construct a dimer model to mimic features of the solid-state geometry. Our simulations predict an average value of 12.7 kJ mol<sup>-1</sup> (no ZPE correction) for the interaction in solid BH<sub>3</sub>NH<sub>3</sub>, thus reassigning it towards the lower end of the dihydrogen bond strength spectrum.

## Computational Methods

A full discussion of the computational methods employed in this work is available in the supplementary information.

## Notes and references

<sup>§</sup>This value at present does not include ZPE correction, which would require the full set of vibrational modes to be averaged over all of k-space. This represents an extremely costly calculation to perform and is currently outwith our computational resources.

<sup>¥</sup>Unfortunately there is currently no experimental sublimation energy available in the literature for comparison.

<sup>⊕</sup>Note, as the PW basis set is by its very nature delocalised, no correction for BSSE is necessary.

<sup>∅</sup>It should be remembered that the current generation of DFT functionals cannot account for non-local dispersion interactions, however such contacts are likely to be of minimal importance for the 1st row element compound BH<sub>3</sub>NH<sub>3</sub>.

[1] T. B. Richardson, S. de Gala, R. H. Crabtree, P. E. M. Siegbahn, *J. Am. Chem. Soc.* 1995, **117**, 12 875.

[2] R. H. Crabtree, P. E. M. Siegbahn, O. Eisenstein, A. L. Rheingold, T. F. Koetzle, *Acc. Chem. Res.* 1996, **29**, 348.

[3] J. Li, F. Zhao, F. Jing, *J. Chem. Phys.* 2002, **116**, 25.

[4] P. L. A. Popelier, *J. Phys. Chem. A* 1998, **102**, 1873.

[5] C. A. Morrison, M. M. Siddick, *Chem. Eur. J.* 2003, **9**, 628.

[6] In this work we take the definitions of lattice and sublimation energies from L. L. Shipman, A. W. Burgess, H. A. Scheraga, *J. Phys. Chem.* 1976, **80**, 52.

[7] W. T. Klooster, T. F. Koetzle, P. E. M. Siegbahn, R. B. Richardson, R. H. Crabtree, *J. Am. Chem. Soc.* 1999, **121**, 6337.

[8] L. R. Thorne, R. D. Suenram, F. J. Lovas, *J. Chem. Phys.* 1983, **78**, 167.